

Flame Test Lab Questions Answer Key:

1. You could readily identify the elements that had obvious colors different from all the others- such as copper that gave off a blue/green color and lithium that gave off a bright red color.
2. It was difficult to identify a couple of the elements that had colors that were similar.
3. It was important to use a different wood splint for each test because the wood splint was contaminated by the prior element which would have effected the color of the element being tested.
4. Each of the compounds are different because they contain a different metal combined with chloride.
5. Each of the compounds was similar in that they were all metal reacted with chloride.
6. The manipulated variable was the element being tested.
7. The responding variable was the color of the flame.
8. The flame was orange at the end of each test because the wood splint started burning.
9. The reason each of the elements produced a different color (or series of wavelengths) is because of the make up of the atom. Each atom consists of a different number of protons, neutrons and electrons that are specific to that type of element. Photons were produced when the energy passed through the atoms and the difference in the atoms (#'s of protons, neutrons and electrons) determines the wavelengths of the photons produced.
10. The production of specific photons by specific elements is very useful to astronomers because they can look at the wavelengths of the photons that are produced by stars to determine the elements that compose that star.

Flame Test Lab Answer Key

K Morrison



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