



LIFE SCIENCES - Grade 10

TOPIC TEST: History of Life on Earth

MARKS: 30

TIME: 30 minutes

SECTION A

QUESTION 1

Various options are given as possible answers to the following questions. Choose the answer and write only the letter (A to D) next to the question number (1.1 to 1.5) in the **ANSWER BOOK**, for example 1.6 D.

- 1.1 Palaeontology is a study of _____
- A. the half-life of radioactive substances.
 - B. continental drift.
 - C. fossils.
 - D. changes during the Cambrian explosion.
- 1.2 Which method allows scientists to say that one fossil is older than the other?
- A. Fossilization
 - B. Relative dating
 - C. Petrification
 - D. Radiometric dating
- 1.3 A period in the earth's history where global temperatures plummet and there is extensive glaciation is known as _____
- A. the greenhouse effect.
 - B. a cold spell.
 - C. a hole in the ozone.
 - D. an ice age.
- 1.4 _____ are the richest sources of fossils.
- A. Layers of strata
 - B. Index fossils
 - C. Mineral salts
 - D. Sedimentary rocks

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Huangqi Zhang



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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web bonus problem here is a balanced net ionic equation $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$ the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web balance each half reaction in basic solution $4\text{S} + 6\text{H}_2\text{O} + 7\text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{NO}_3^- + \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox

reaction in acid solution using the half reaction method
8 H_2O_2 $\text{Cr}_2\text{O}_7^{2-}$ 9 TeO_3 2 N_2O_4 O_2 Cr^{3+} TeO_4 11 PbO_2 IO_2 Pb^{2+} 12
as NO_3^- IO_3^- H_2AsO_4^-

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H^+ OH^- and e^- in this order balance the atoms in the equation apart from O and H

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ 2HNO_2 $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}_2$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2^- $\text{C}_6\text{H}_8\text{O}_6$ and NO_2^- 2 and $\text{C}_6\text{H}_8\text{O}_6$ NO_2^- and $\text{C}_6\text{H}_8\text{O}_6$ H^+ and NO_2^- $\text{C}_6\text{H}_8\text{O}_6$ and NO_2^-