

The Tangent Ratio



DIAGRAMS NOT TO SCALE

Section A

Write down the value of $\tan \theta$ for each triangle.
Do not use a calculator, leave your answer as a fraction.



$\tan \theta =$



$\tan \theta =$



$\tan \theta =$



$\tan \theta =$

Section B

Calculating Angles

1) Find the value of θ in each equation. Give your answer correct to 2.d.p.

a) $\tan \theta = 0.54$
 $\theta =$

b) $\tan \theta = 0.185$
 $\theta =$

c) $\tan \theta = \frac{8}{13}$
 $\theta =$

d) $\tan \theta = 4.15$
 $\theta =$

2) Find the size of each angle labelled θ . Give your answer correct to 2.d.p.



$\theta =$



$\theta =$



$\theta =$



$\theta =$



$\theta =$



$\theta =$

Section C

Calculating side lengths

1) Use the tangent ratio to find length of the opposite side in each right angled triangle.
Give your answer correct to 2.d.p.



$x =$



$x =$



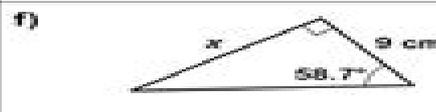
opposite =



$x =$



$x =$



$x =$

Practice 8 2 The Tangent Ratio Answer Key

Paul Steenson



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Scott Foresman-Addison Wesley Middle School Math, 2002

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Algebra One-[two] Glen D. Vannatta, 1962

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web thermochemical equations equations that represent both mass and enthalpy changes h 2 o s h 2 o l h 6 01 kj mol this is an endothermic process it requires 6 01 kj to melt one mole of ice h 2 o s the enthalpy value will change if the number of moles varies from the 1 1 reaction stoichiometry this is an exothermic process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy e of a system in any process is equal to the heat q added to the system plus the work w done on the system by its surroundings e

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