

General Chemistry II Worksheet
Chapters 18-19

1) Calculate the pH of a solution that is 0.322 M in nitrous acid (HNO₂) and 0.178 M in potassium nitrite (KNO₂).

The acid dissociation constant of nitrous acid is 10^{-4}
4.50 ×

- A) 3.093 B) 4.554 C) 3.607 D) 10.91 E) 14.26

2) Which of the following acids (listed with K_a values) and their conjugate base should be used to form a buffer with a pH of 2.34?

- A) C₆H₅OH, K_a = 1.3×10^{-10}
B) HCN, K_a = 4.9×10^{-10}
C) HClO₂, K_a = 1.1×10^{-2}
D) HIO, K_a = 2.3×10^{-10}
E) HN₃, K_a = 2.5×10^{-5}

3) What is the pH of the resulting solution if 45.00 mL of 0.10 M acetic acid is added to 10.00 mL of 0.10 M KOH? Assume that the volumes of the solutions are additive. K_a = 1.8×10^{-5} for

CH₃CO₂H:

- A) 9.80 B) 8.71 C) 6.62 D) 5.29 E) 4.20

What is the molar solubility of silver sulfate (Ag₂SO₄) in water? The solubility-product

4) constant for Ag₂SO₄

is 1.5×10^{-5} at

- A) 5×10^{-3} mol/L
B) 5×10^{-4} mol/L
C) 6×10^{-2} mol/L
D) 7×10^{-2} mol/L
E) 5×10^{-2} mol/L

In which of the following solutions would solid PbCl₂ be expected to be least soluble at 25°C?

5) the

- A) 0.1 M NaCl B) 0.1 M KNO₃ C) 0.1 M HCl D) 0.1 M CaCl₂

6) A solution contains 2.2×10^{-3} M in Cu²⁺ and 0.33 M in LiCN. If the K_f for Cu(CN)₄²⁻ is 1.0×10^{25} , how much copper ion remains at equilibrium?

- A) 1.9×10^{-26} M
B) 6.7×10^{-28} M
C) 2.9×10^{-27} M
D) 4.6×10^{-25} M
E) 3.8×10^{-24} M

7) Which of the following processes has a ΔS_{sys} > 0?

- A) potassium bromide dissolves in pure water
B) $2 \text{NH}_3(g) + \text{CO}_2(g) \rightarrow \text{NH}_2\text{CONH}_2(aq) + \text{H}_2\text{O}(l)$
C) $2 \text{HBr}(g) \rightarrow \text{H}_2(g) + \text{Br}_2(l)$

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