

Gas Laws – Review Problems

Instructions: Use the Combined Gas Law and Dalton's Law to complete each problem.

Boyle's Law:

$$P_1 V_1 = P_2 V_2$$

Charles' Law:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Gay-Lussac's Law:

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- 1) If the gas inside a volleyball has a pressure of 5.50 atm at a temperature of 309 K, what will the pressure be if the volleyball is reaches a temperature of 279 K?

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \rightarrow \quad \frac{5.50}{309} = \frac{P_2}{279} \quad \rightarrow \quad \boxed{P_2 = 4.97 \text{ atm}}$$

- 2) An inflated balloon contains gas with a volume of 2.5 liters until the pressure on the balloon is changed, increasing the volume of the balloon to 4.7 liters. If the final pressure of the balloon is 2.6 atm, what was the initial pressure?

$$P_1 V_1 = P_2 V_2 \quad P_1 (2.5) = (2.6)(4.7) \quad \boxed{P_1 = 4.89 \text{ atm}}$$

- 3) A car is left outside at night and has a tire pressure of 2.08 atm. In the morning, the temperature is 282 K and the tire pressure has increased to 2.29 atm. What was the initial temperature the previous night?

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \rightarrow \quad \frac{2.08}{T_1} = \frac{2.29}{282} \quad \rightarrow \quad \boxed{T_1 = 256.14 \text{ K}}$$

- 4) A canister is filled with 5.75 liters of gas and heated to a temperature of 600 K. Afterwards, the gas cools to a temperature of 340 K. What was the final volume?

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \rightarrow \quad \frac{5.75}{600} = \frac{V_2}{340} \quad \rightarrow \quad \boxed{V_2 = 3.26 \text{ L}}$$

- 5) A large balloon is filled with 4.65 L of gas at a pressure of 1.21 atm. It is then placed under pressure until the volume decreases to 2.22 L. How much pressure was used?

$$P_1 V_1 = P_2 V_2 \quad (1.21)(4.65) = P_2 (2.22) \quad \boxed{P_2 = 2.53 \text{ atm}}$$

Gas Law Study Guide Answers

BM King



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