

10 • Gas Laws

UNIT REVIEW

Directions: Use the following questions to guide your studying. Also, review your notes packet and look at questions you missed on quizzes and homework. Doing this study guide does not guarantee that you will ace the test. You need to ask questions if you have them!

Useful Information

STP = 0°C = 273 K and
 1 atm = 760 torr = 760 mmHg
 = 101.3 kPa = 14.7 psi
 = 14.7 $\frac{\text{lb}}{\text{in}^2}$

Ideal Gas Constant, R
 = 62.4 $\frac{\text{L}\cdot\text{mmHg}}{\text{mol}\cdot\text{K}}$
 = 0.0821 $\frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$
 = 8.31 $\frac{\text{L}\cdot\text{kPa}}{\text{mol}\cdot\text{K}}$

1. Perform the following conversions.

a. $26.0^\circ\text{C} = \underline{299 \text{ K}}$

$$26.0^\circ\text{C} + 273 = 299$$

b. $135^\circ\text{C} = \underline{408 \text{ K}}$

$$135^\circ\text{C} + 273 = 408$$

c. $-127^\circ\text{C} = \underline{146 \text{ K}}$

$$-127^\circ\text{C} + 273 = 146 \text{ K}$$

d. $100 \text{ K} = \underline{-173}^\circ\text{C}$

$$100 \text{ K} = x + 273, \quad x = -173$$

e. $400 \text{ K} = \underline{127}^\circ\text{C}$

$$400 \text{ K} = x + 273$$

$$x = 127$$

f. $4 \text{ K} = \underline{-269}^\circ\text{C}$

$$4 = x + 273$$

$$x = -269$$

2. What is the temperature of a sample of gas that has double the kinetic energy (energy of motion) of a sample at 80 °C.

$$80^\circ\text{C} = 353 \text{ K}$$

Kelvin temperature
 measures kinetic energy

If you double the
 kinetic energy, then
 you double the Kelvins.

$$2(353 \text{ K}) = 706 \text{ K}$$

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