

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- | | |
|-----------------------------|---|
| a. Sulfur hexafluoride | SF_6 |
| b. Lithium nitride | Li_3N |
| c. Chromium (III) carbonate | $\text{Cr}_2(\text{CO}_3)_3$ |
| d. Tin (II) chloride | SnCl_2 |
| e. Ammonium acetate | $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$) |
| f. Mercury (I) chloride | Hg_2Cl_2 |
| g. Potassium bromate | KBrO_3 |
| h. Hydrosulfuric acid | $\text{H}_2\text{S(aq)}$ |
| i. Chloric acid | HClO_3 |
| j. Sulfurous acid | H_2SO_3 |

Name each of the following compounds:

- | | |
|------------------------------|-------------------------|
| a. CuSO_4 | Copper (II) sulfate |
| b. AlF_3 | aluminum fluoride |
| c. HI | hydrogen iodide |
| d. NO | nitrogen monoxide |
| e. $\text{H}_2\text{Se(aq)}$ | hydroselenic acid |
| f. HNO_3 | nitric acid |
| g. HNO_2 | nitrous acid |
| h. NaHSO_4 | sodium hydrogen sulfate |

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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Ensheng Dong



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Table of Contents Chemistry Chapter 10 Study Guide Answer Key

1. Understanding the eBook Chemistry Chapter 10 Study Guide Answer Key
 - The Rise of Digital Reading Chemistry Chapter 10 Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 10 Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 10 Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 10 Study Guide Answer Key
 - Personalized Recommendations
 - Chemistry Chapter 10 Study Guide Answer Key User Reviews and Ratings
 - Chemistry Chapter 10 Study Guide Answer Key and Bestseller Lists

5. Accessing Chemistry Chapter 10 Study Guide Answer Key Free and Paid eBooks
 - Chemistry Chapter 10 Study Guide Answer Key Public Domain eBooks
 - Chemistry Chapter 10 Study Guide Answer Key eBook Subscription Services
 - Chemistry Chapter 10 Study Guide Answer Key Budget-Friendly Options
6. Navigating Chemistry Chapter 10 Study Guide Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Chapter 10 Study Guide Answer Key Compatibility with Devices
 - Chemistry Chapter 10 Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 10 Study Guide Answer Key
 - Highlighting and Note-Taking Chemistry Chapter 10 Study Guide Answer Key
 - Interactive Elements Chemistry Chapter 10 Study Guide Answer Key
8. Staying Engaged with Chemistry Chapter 10 Study Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 10 Study Guide Answer Key
9. Balancing eBooks and Physical Books Chemistry Chapter 10 Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 10 Study Guide Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 10 Study Guide Answer Key
 - Setting Reading Goals Chemistry Chapter 10 Study Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 10 Study Guide Answer Key
 - Fact-Checking eBook Content of Chemistry Chapter 10 Study Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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