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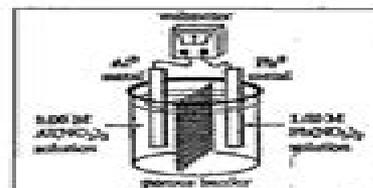
Ch. 17 – Electrochemistry: Practice Problems I

- 1) The half-reaction that occurs at the cathode during the electrolysis of molten sodium bromide is ____.
- a) $2\text{Br}^- \rightarrow \text{Br}_2 + 2\text{e}^-$
 b) $\text{Br}_2 + 2\text{e}^- \rightarrow 2\text{Br}^-$
 c) $\text{Na}^+ + \text{e}^- \rightarrow \text{Na}$
 d) $\text{Na} \rightarrow \text{Na}^+ + \text{e}^-$
- NaBr*
 ↑ reduction rxn. $\text{Na}^+ + \text{e}^- \rightarrow \text{Na}$
 Anode: Oxidation $2\text{Br}^- \rightarrow \text{Br}_2 + 2\text{e}^-$

- 2) What voltage will be produced by this electrochemical cell?

- a) +2.97V
 b) +1.55V $E_{\text{cell}} = E_{\text{cathode}} - E_{\text{anode}}$
 c) -1.81V $E_{\text{cell}} = (-0.13) - (-1.68)$
 d) -2.97V $E_{\text{cell}} = +1.55\text{V}$

Reduction Potential, E°	
$\text{Pb}^{2+} + 2\text{e}^- \rightarrow \text{Pb}$	-0.13 V
$\text{Al}^{3+} + 3\text{e}^- \rightarrow \text{Al}$	-1.68 V



more negative $E^\circ \rightarrow$ oxidation

- 3) A galvanic cell is set up using the following half-reactions:

$\text{Zn} | \text{Zn}^{2+} (1\text{M}) || \text{Ni}^{2+} (1\text{M}) | \text{Ni}$
 Anode Cathode
 What is the standard voltage of the cell?

$\text{Ni}^{2+} + 2\text{e}^- \rightarrow \text{Ni}$	$E^\circ = -0.250\text{V}$
$\text{Zn}^{2+} + 2\text{e}^- \rightarrow \text{Zn}$	$E^\circ = -0.763\text{V}$

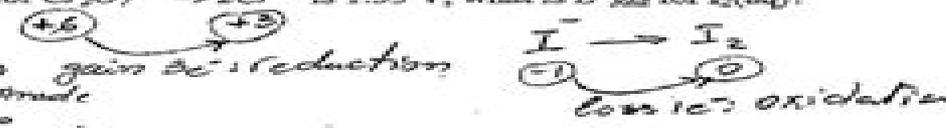
- a) -0.513 V b) -1.013 V c) +1.013 V d) +0.513 V $E_{\text{cell}} = (-0.250) - (-0.763)$
 $E_{\text{cell}} = +0.513\text{V}$

- 4) For the reaction shown below, $E^\circ_{\text{cell}} = 0.79\text{V}$.



Given that the standard reduction potential for $\text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+}$ is 1.33 V, what is E°_{red} for $\text{I}_2(\text{aq})$?

- a) +0.54 V
 b) -0.54 V
 c) +0.18 V
 d) -0.18 V
- $E_{\text{cell}} = E_{\text{cathode}} - E_{\text{anode}}$
 $+0.79 = (1.33) - E_{\text{anode}}$
 $-0.54 = -E_{\text{anode}} \rightarrow E_{\text{anode}} = +0.54\text{V}$

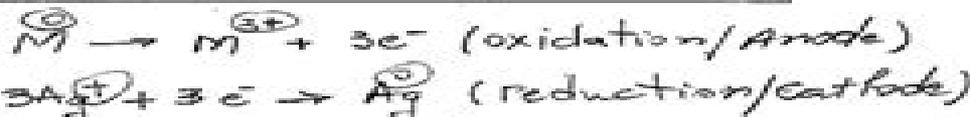


- 5) According to the information below, what is the standard reduction potential for the half-reaction



- a) -1.66 V
 b) -0.06 V
 c) 0.06 V
 d) 1.66 V
 e) 3.26 V

$\text{M} (\text{s}) + 3\text{Ag}^+ (\text{aq}) \rightarrow 3\text{Ag} (\text{s}) + \text{M}^{3+} (\text{aq})$	$E_{\text{cell}}^\circ = +2.46\text{V}$
$\text{Ag}^+ (\text{aq}) + \text{e}^- \rightarrow \text{Ag} (\text{s})$	$E^\circ = +0.80\text{V}$



$E_{\text{cell}}^\circ = E_{\text{cathode}}^\circ - E_{\text{anode}}^\circ$
 $+2.46 = (+0.80) - E_{\text{anode}}^\circ$
 $E_{\text{anode}}^\circ = -1.66\text{V}$

Electrochemistry Practice Problems

SB Merriam

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