

Unit 1: Atomic Structure and Theory Test Review

Directions: using our notes from this unit complete the following study guide. You may not use the guide on your test but I will be a good tool to help you study

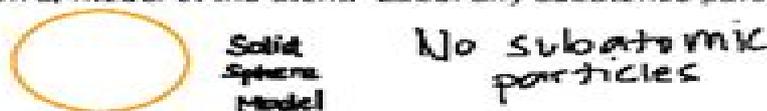
1. Know the definition of the following vocabulary words:

- Proton- **positively charged particle found in the nucleus**
- Nucleus- **dense positive center of the atom**
- Electron- **negatively charged particle found outside the nucleus**
- Atomic number- **gives atom identity, matches the # of protons**
- Atom- **what all matter is made of**
- Neutron- **neutral subatomic particle found in the nucleus**

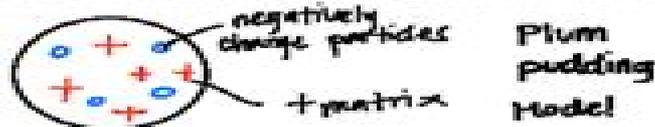
2. Answer the following questions about the development of the atomic model:

- Name person first proposed the idea of the atom? Democritus
- Name the scientist that performed the gold foil experiment? Rutherford
- What was John Dalton's model of the atom called? Solid Sphere Model
- What did JJ Thomson discover about the atom after his experiment?
Electrons
- Name three pieces of information Ernest Rutherford discover about the atom after his experiment. Atoms are mostly empty space The atom has a small, dense nucleus.
The nucleus is tiny compared to the atom as a whole
- What happened to some of the alpha particles when they hit the gold foil during Rutherford's experiment?
Most passed straight through the foil
Some were deflected at small angles
A few bounced back
- What happened to the cathode ray beam when it was exposed to a magnet?
The beam deflect toward the + portion of the magnet

h. Draw John Dalton's, model of the atom. Label any subatomic particles their atom contains if there are any.



i. Draw JJ Thomson's, model of the atom. Label any subatomic particles their atom contains if there are any.



Electron Structure Test Review Guide

VM Jensen



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