

Chapter 20 Electrode potentials and electrochemical cells

Practice questions

Question number	Answer	Marks	Guidance
1 (a)	By definition	1	Allow 'set to this value'
1 (b)	1.23V	1	Allow + or -
1 (c)	$\text{Pt} \text{H}_2(\text{g}) \text{OH}^-(\text{aq}),\text{H}_2\text{O}(\text{l}) \text{O}_2(\text{g}) \text{H}_2\text{O}(\text{l}),\text{OH}^-(\text{aq}) \text{Pt}$ Correct but with Pt missing Includes Pt with correct representation	1 1	H_2O not essential, allow reverse order
1 (d)	Uses $\text{O}_2 + 2\text{H}_2\text{O} + 4\text{e}^- \rightarrow 4\text{OH}^-$ And (2x) $2\text{OH}^- + \text{H}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{e}^-$ $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	1 1	
1 (e)	Increases the surface area (so reaction faster)	1	
1 (f)	Overall reaction is the same ($2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$)	1	Or shows e.m.f. is the same
1 (g)	Hydrogen and oxygen supplied continuously OR Can be operated without stopping to recharge	1	Or can be refuelled quickly Allow any one mark
1 (h)	Hydrogen may need to be made using an energy source that is not 'carbon neutral'	1	
2 (a)	Zn	1	This is always the species with the most negative E value (it is oxidised itself).
2 (b) (i)	Fe^{3+}	1	
2 (b) (ii)	Cl_2	1	
2 (c) (i)	The standard electrode potential: 1.25V	1	Remember that: $E_{\text{cell}} = E_{\text{right}} - E_{\text{left}}$
2 (c) (ii)	$\text{Ti}^{3+} + 2\text{Fe}^{2+} \rightarrow 2\text{Fe}^{3+} + \text{Ti}^+$	2	One mark for a balanced equation and one for the correct direction.
3 (a)	$\text{Pt} \text{H}_2 \text{H}^+ \text{Fe}^{3+} \text{Fe}$ Note, allow one mark only for correct symbol in reverse: $\text{Fe} \text{Fe}^{2+} \text{H}^+ \text{H}_2 \text{Pt}$	2	Allow 1 for correct order of symbols but lose 2nd mark for a wrong phase boundary(s) / Pt missing / extra Pt on RHS, additional phase boundary Allow dashed lines for salt bridge Ignore state symbols Ignore 2 if used before H^+
3 (b)	Electron donor	1	Allow (species that) loses electrons Do not allow reference to electron pairs
3 (c)	Cl_2 / chlorine (Species on RHS / electron donor) has most positive / largest E_0 / has highest potential	1 1	If M1 blank or incorrect cannot score M2 Do not allow reference to e.m.f. or $E(\text{cell})$
3 (d) (i)	Cl / chlorine	1	
3 (d) (ii)	Chlorine +1 to chlorine 0	1	CE if chlorine not identified in 3(d)(i) Allow chlorine +1 to chlorine -1 (in Cl^-) Allow oxidation state decreases by one / two Allow oxidation state changes by -1 OR -2
3 (e)	$4\text{HOCl} + 4\text{H}^+ + 4\text{OH}^- \rightarrow 2\text{Cl}_2 + \text{O}_2 + 6\text{H}_2\text{O}$ OR $4\text{HOCl} \rightarrow 2\text{Cl}_2 + \text{O}_2 + 2\text{H}_2\text{O}$	2	Allow one mark for any incorrect equation that shows $\text{HOCl} \rightarrow \text{Cl}_2 + \text{O}_2$ Allow multiples Ignore state symbols Penalise one mark for uncanceled / uncombined species (eg $\text{H}_2\text{O} + \text{H}_2\text{O}$ instead of $2\text{H}_2\text{O}$)
3 (f)	$\text{e.m.f.} = 0.40 - (-1.25) = 1.65\text{V} / +1.65\text{V}$	1	Allow -1.65V

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Michael Rajnik



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in either case $f(x) = 0$ for $x \in [p, p + \delta]$ since f is continuous at p and $f(p) = 0$ there exists a $\delta > 0$ with $f(x) = 0$ for $x \in [p, p + \delta]$ and a $x \in [p, p + \delta]$ we restrict δ so that $p \in [p, p + \delta]$ is a see more

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d the relative error in part b is 0 the relative error in part c is 0 1 a $\lim_{x \rightarrow 0} \frac{e^x - e^{-x}}{x} = \lim_{x \rightarrow 0} \frac{e^x + e^{-x}}{1}$ see more

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by the intermediate value theorem 1 applied to the interval with endpoints x_1 and x_2 there exists a number ξ between x_1 and x_2 for which $f(\xi) = \frac{f(x_2) - f(x_1)}{x_2 - x_1}$ see more

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so the actual error is $2 \cdot 10^{-7} \cdot 16$ a $p_3(x) = 3x^3 - 16x^2 + 23 \cdot 648x^3$ b we have $f(4) = 119 \cdot 1296 \cdot \sin x$ 3 see more